

Name _____ Classroom Number _____ Class Period _____

Periodic Table Practice

Using your Periodic Table of the Elements identify the following for each Element.

Element Name	Element Symbol	Atomic #	Atomic Mass	# Protons	# Neutrons	# Electrons
		1				
	Au					
			207			
Oxygen						
				11		
	Ne					
						23
Iron						
			59			
		12				
				29		
						5
					51	40
Platinum						
		36				
			184			
	N					
					18	17
Argon						
	Xe					

Continued Thinking....

1. Elements are considered "stable" when their outer electron shell is full. Below is an illustration showing how many electrons fill each level of an element. Using this diagram and your evidence gathered about each element, identify those that are "stable".

Valence Electrons

- Each electron shell can hold a certain number of electrons
- Electron shells are filled from the **inside** out

Electron Shell	Number of Electrons
1	2
2	8
3	8
4	18
5	18
6	32
7	32

- 1.
- 2.
- 3.
- 4.

2. If you know that an atom has 6 electrons on the second energy level and no electrons on the third energy level, explain how you know that this atom must be oxygen?
3. IF, a newly developed microscope discovered that the electron configurations for each shell was inaccurate and the first shell of an atom actually held 10 electrons rather than 2 how would this effect the list above? What element or elements would now be considered "stable"?
4. The area of the atom where the protons and neutrons are found is called the _____.
5. The atomic number of an atom is equal to the number of _____.
6. The _____ equal the number of protons in the element.
7. A _____ has no charge, a _____ has a positive charge and an _____ has a negative charge.